

YASNETSOV, V. S.

USSR / Human and Animal Physiology (Normal and Pathological). The Sense Organs. Vision. T

Abs Jour: Ref Zhur-Biologiya, No 21, 1958, 97986

Author : Popov. M. Z.; Yasnetsov, V. S.

Inst : Smolensk Medical Institute

Title : The Influence of Aminazine on Intraocular and Blood Pressure in Glaucoma Patients

Orig Pub: Trudy Smolenskogo med. in-ta, 1957, 6, 297-306

Abstract: no abstract

Card 1/1

YASNETSOV, V.S.

EXCERPTA MEDICA Sec.2 Vol.11/5 Physiology, etc. May 58

2287. EFFECT OF CHLORPROMAZINE ON GASTRIC SECRETION (Russian text):
Yasnetsov V. S. Dept. of Pharmacol., Smolensk Med. Inst., Smolensk -
FARMAKOL. I TOKSIKOL. 1957, 20/1 suppl. (33)

136 experiments were performed on 7 dogs with the small Pavlov stomach. After i. m. injection of 5 mg./kg. of chlorpromazine there was marked inhibition of secretion, both free and bound HCl were decreased, and the enzymatic potency of the secretion dropped for 3-4 hr., then rose to far above the normal values. Chlorpromazine caused prolonged sleepiness and weakness with occasional periods of restlessness. Doses of 2 mg. and 10 mg. per kg. gave similar effects.

Vacek - Brno

~~YASNETSOV V.S.~~

Effect of dimedrol on gastric secretion in dogs. Farm. i toks.
20 no.6:52-54 N-D '57 (MIRA 11:6)

1. Kafedra farmakologii Smolenskogo meditsinskogo instituta
(nauchnyy rukovoditel' raboty - prof. G.A. Ponomarev).
(GASTRIC JUICE,
secretion, eff. of diphenhydramine in dogs (Rus))
(DIPHENHYDRAMINE, effects.
on gastric secretion in dogs (Rus))

YASNETSOV, V. S., Candidate Med Sci (diss) -- "The effect of phenothiazine derivatives (aminazine, diprazine, and ethizine) and demedrol on the secretory function of the stomach (Experimental investigation)". Smolensk, 1959. 19 pp (Minsk State Med Inst), 250 copies (KL, No 23, 1959, 174)

YASNETSOV, V.S., dotsent

Central mechanisms regulating intraocular pressure. Trudy SMI 16:206-212 '63. (MIRA 18:1)

Effect of some pharmacological agents on the activity of carbonic anhydrase in the tissues of the eye. Ibid.:213-216

Elastotonometric studies during the influence of phenothiazine derivatives. Ibid.:240-245

1. Iz kafedry farmakologii (zav. - prof. A.I.Mitrofanov) i kafedry glaznykh bolezney (zav. - prof. M.Z.Popov) Smolenskogo gosudarstvennogo meditsinskogo instituta.

POPOV, M.Z., prof.; YASNETSOV, V.S., dotsent

Aminazine and diphacil in treating glaucoma. Trudy SMI 16:217-227 '63.
(MIRA 18:1)

1. Iz kafedry glaznykh bolezney (zav. - prof. M.Z.Popov) i kafedry
farmakologii (zav. - prof. A.I.Mitrofanov). Smolenskogo gosudarstvennogo
meditsinskogo instituta.

YASNETSOV, V.S., dotsent; YASNETSOVA, N.M., kand. med. nauk

Fonurit and aminazine in the treatment of glaucoma. Trudy SMI 16:228-239 '63. (MIRA 18:1)

Elastotonometric studies during the influence of urea. Ibid.:246-252

1. Iz kafedry farmakologii (zav. - prof. A.I.Mitrofanov) i kafedry glaznykh bolezney (zav. - prof. M.Z.Popov) Smolenskogo gosudarstvennogo meditsinskogo instituta.

YASNETSOVA N.M.
USSR/Human and Animal Morphology - Normal and Pathological.
Organs of the Senses.

S

Abs Jour : Ref Zhur Biol., No 23, 1958, 106018

Author : Yasnetsova, N.M.

Inst :

Title : Pathohistological Changes after Anterior Iridosclerotomy.
(Experimental Histological Investigation)

Orig Pub : Zdravookhr. Belorussii, 1958, No 6, 22-25

Abstract : The technique of anterior iridosclerotomy is described.
Operations on dogs showed the dynamics of the regenerative process with formation of the infiltrating scar - one of the elements upon which the hypotensive effect of the operation is conditioned.

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YASNETSOVA, N.A., kand. med. nauk

Diathermocoagulation of retinal detachment. Trudy SMI 17:114-117
'63.

Comparative results of diathermocoagulation and drug therapy of
retinal detachment. Ibid.:118-121

(MIRA 18:1)

1. Iz kafedry glaznykh bolezney (zav. - prof. M.Z. Popov) Smolenskogo
gosudarstvennogo meditsinskogo instituta.

YASNETSOV, V.S., dotsent; YASNETSOVA, N.M., kand. med. nauk

Fonurit and aminazine in the treatment of glaucoma. Trudy SMI 16:228-239 '63. (MIRA 18:1)

Elastotonometric studies during the influence of urea. Ibid.:246-252

1. Iz kafedry farmakologii (zav. - prof. A.I.Mitrofanov) i kafedry glaznykh bolezney (zav. - prof. M.Z.Popov) Smolenskogo gosudarstvennogo meditsinskogo instituta.

YASNEV, D.H.

~~BERKNOV~~, Dmitriy Ivanovich; YASNEV, D.A. [deceased], red.; TSYMBALIST, N.N.,
red. izd-va; ZMF, Ye.M., tekhn red.

[Crushing equipment in concentration and crushing plants] Drobil'-
noe oborudovanie obogatitel'nykh i drobil'nykh fabrik. Sverdlovsk,
Gos. nauchno-tekhn. izd-vo lit-ry po chernoi i tsvetnoi metallurgii,
Sverdlovskoe otd-nie, 1958. 295 p. (MIRA 11:7)
(Crushing machinery)

YASNEV, Yu.

Collapse of the colonial system in Africa. Komm.Voeruzh.Sil 1
no.2:35-39 Ja '61. (MIRA 14:8)
(Africa--Politics) (Colonies)

YASNEV, Yu.

Aggressors in jungles. Av. i kosm. 45 no.11:92-95 '62.
(MIRA 15:11)
(Vietnam, South—Military assistance, American)

SOV/86-58-7-37/38

AUTHOR: Yasnev, Yu. N., Capt

TITLE: Playing With Fire (Opasnaya igra s ognem)

PERIODICAL: Vestnik vozdushnogo flota, 1958, Nr 7, pp 90-95 (USSR)

ABSTRACT: This is a propaganda article dealing with flights carried out by the Strategic Air Command of the USA in the direction of and along the Soviet borders.

Card 1/1

YASNEV, Yu.N., mayor

Foci of aerial espionage and provocations. Vest.Vozd.Fl. no.9:
89-92 S'60. (MIRA 13:10)
(Military bases, American)

YASNEVA, A.

Sowing in spring? no, earlier. IUn.tekh. 3 no.3:57-58
Mr '59. (MIRA 12:4)

(Sowing)

YASNEVA, Ye. Ye.

Coast changes of the Kuybyshev Reservoir. Vop. gidrogeol. i
inzh. geol. no.20:149-155 '62. (MIRA 16:4)

(Kuybyshev Reservoir—Coast changes)

YASNIKOV, A. A. (Co-author)

See: SHILOV, Ye. A.

Yasnikov, A. A. and Shilov, Ye. A. - "The benzoin condensation of aldehydes with groups of type II", (Report), Soobshch. o nauch. rabotakh chlenov Vsesoyuz. khim. o-va im. Mendeleyeva, 1949, Issue 2, p. 8-9.

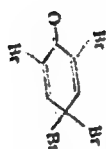
SO: U-4630, 16 Sept. 53, (Letopis 'Zhurnal 'nykh Statey, No. 23, 1949).

B7I-6

BA

"Danger zone" of chlorine bleaching. E. A. Shilov and A. A. Yasnikov. (Tselul. i pap., 1950, No. 11, 35-38).—The "danger zone," in which max. fibre damage occurs in hypochlorite bleaching, is in the pH range 6-8. The rates of oxidation of acids and D-glucosonic acids, starch, EtOH, and cellulose by HOCl all show a sharp max. at pH 6-8. By analogy with the mechanism of oxidation of saturated alcohols by HOCl, it is supposed that hydroxy-acids, polyhydric alcohols, and cellulose are oxidized with the intermediate formation of unstable hypochlorite esters in 2 stages: (1) $\text{CH-OH} + \text{HOCl} \rightarrow \text{CH-OCl} + \text{H}_2\text{O}$; (2) $\text{CH-OCl} + \text{OH}^- \rightarrow \text{CO} + \text{H}_2\text{O} + \text{Cl}^-$. At low pH, when HOCl is present and $[\text{OH}^-]$ is low, the reaction rate is determined by stage (2) and increases approx. $\propto [\text{OH}^-]$; in the alkaline region it is mainly determined by stage (1), and depends on $[\text{HOCl}]$; thus it reaches a max. at pH ~7. Bleaching proper, i.e., the destruction of coloured compounds, proceeds by a different mechanism and the rate falls with increase of pH. From the data of Chibbens and Ridge (B., 1937, 471) on the effect of pH on cellulose oxidation it is concluded that the range pH 8-9-10 avoids the "danger zone" but maintains an adequate rate of bleaching. The range pH 2-3.5 also avoids the danger zone and gives a max. bleaching rate; this suggests the practical possibility of chlorine-water bleaching. E. B. Uvarov.

The structure of tribromophenol bromide and some data on the theory of its formation. A. A. Yankov and R. A. Sidor. Doklady Akad. Nauk S.S.S.R. 76 (1959, 1961). - The literature is reviewed (7 references). An aq. suspension of 2,4,6-Br₃C₆H₂OH (I) was mixed with radioactive Br water (active addn. divided; 1 part was reduced with aq. Na₂SO₃, filter paper and their activity detd. The fraction of residual activity in making the tribromophenol bromide (II). If the suspension of I is added to Br water, then upon mixing the suspension of I is added to Br water, then upon reduction of all 3 Br atoms; an increase of residual activity is retained, but if the addn. of the activity, the product on reduction retains only 25% of the activity, indicating exchange of all 3 Br atoms. II in aq. media, sun does not exchange Br with Br water or NaBr, but slow exchange occurs with concd. HBr in aq. or CCl₄ media. The exchange of Br between I and II is very slow in aq. solution, but is accelerated by a transfer agent like pyridine. Apparently the "active" state of Br during the reaction is explained by the structure of II. The above results show that II has the structure of a quinone bromide, and not that of a hypobromite. I does not react with Br in CCl₄, or A-OH, but a II acceptor (pyridine or H₂O) gives an immediate ppt. of II; probably the II is detached from the H₂O of the phenol and the product assumes the final structure:



G. M. Kosolapoff

Inst. Org. Chem., Ukr. SSR

YASNIKOV, A. A.

USSR/Chemistry - Chlorination 21 Nov 51

"Kinetics of the Attachment of Hypochlorous Acid to Molecules Having a Double Bond of Low Activity," Ye. A. Shilov, Active Mem, Acad Sci Ukrainian, SSR G. V. Kupinskaya, A. A. Yasnikov, Inst of Org Chem, Acad Sci Ukrainian SSR "Dok Ak Nauk SSSR" Vol LXXXI, No 3, pp 435-438

Trimethylallylammonium ion was chosen as the compd for this work because of its soly in water and the low activity of its double bond. In the absence of HCl, HOCl combines with trimethylallylammonium perchlorate according to the eq - $\frac{dC}{dt} = k_3 \frac{A}{[HOCl]}^2$ where A is concn of trimethylallylammonium perchlorate, C is concn of active chlorine in moles per liter. $k_3 = 3.5$ at 25° and 1.7 at 15° . Cl_2O is shown to be the active agent of the reaction.

PA 214T21

YASNIKOV, A. A.

Kinetics and mechanism of reactions of active chlorine with organic compounds IX Oxidation of glycolaldehyde and glyceraldehyde E. A. Shklovskiy and A. A. Yasnikov

At high pH the enolate form appears to be oxidized. X. Oxidation of hydroxy acids. Ibid. 611-24. Oxidation of lactic acid (I) and D-gluconic acid (II) as well as preliminary study of pyruvic acid (III). The rate of oxidation of I

Value of rate constants with increase of pH is as follows: for I and 4 $\times 10^{-4}$ for II time in min. At pH 10 the rate of oxidation of I and II is given best by the dependence where k_1 for I goes from 7.5×10^{-4} at pH 1 to 1.5×10^{-4} at pH 10, for II these are 2.5×10^{-4} and 4×10^{-4} at pH 1 and 10 respectively. In medium the active agent is proceeding through two steps:

At pH 10 the enolate form appears to be oxidized. X. Oxidation of hydroxy acids. Ibid. 611-24. Oxidation of lactic acid (I) and D-gluconic acid (II) as well as preliminary study of pyruvic acid (III). The rate of oxidation of I

1. SHILOV, ACAD. YE. A; YASNIKOV, A. A.
2. USSR (600)
4. Amino Acids
7. Kinetics and mechanism of the enolization of acetone in the presence of amino acids. Acad. E. A. Shilov, A. S. Yasnikov. Dokl. AN SSSR 84 No. 2, 1952. Institut Organicheskoy Khimii Akademii Nauk Ussr. Red. 17 March 1952.
9. Monthly List of Russian Accessions, Library of Congress, September 1952. UNCLASSIFIED.

YASNIKOV, A. A.

USSR/Chemistry - Physical chemistry

Card 1/1 Pub. 147 - 22/27

Authors : Shilov, E. A., Yasnikov, A. A.

Title : Regarding the structure of isomeric pentachlorophenol chlorides

Periodical : Zhur. fiz. khim. 28/9, 1680-1681, Sep 1954

Abstract : Comments, on the report by P. P. Shorygin and M. I. Gostev stating that low-melting isomer of pentachlorophenol chloride has a structure similar to that of hypochlorous acid esters, are presented. The use of marked Cl, for the purpose of determining the structure of isomeric pentachlorophenol chlorides, is recommended. Four USSR references (1950-1954).

Institution : ...

Submitted : June 14, 1954

Yasnikov, A. A.

USSR/ Chemistry - Organic chemistry

Card 1/1 Pub. 116 - 7/29

Authors : Matkovskiy, K. I., and Yasnikov, A. A.

Title : Synthesis of levulinic acid $1C^{14}$

Periodical : Ukr. khim. zhur. 21/6, 721-722, Dec 1955

Abstract : A new method is described for the synthesis of levulinic acid - $1C^{14}$ from $NaC^{14}N$ and methyl-beta-chlorethylketone. Saponification of the obtained nitrile in hydrochloric acid led to the formation of levulinic acid - $1C^{14}$. The effect of temperature on the yield of levulinic acid nitrile is explained. Three references: 1 USA, 1 Germ. and 1 Fr. (1936-1949).

Institution : Acad. of Sc. Ukr. SSR, Inst. of Organ. Chem.

Submitted : June 13, 1955

and extd. with H₂O and HCl. The ext. was
evaporated and pptd. NaCl

and K. L. Murty
the amt. of labeled carotene in the leaves was
labeled levulinic acid, pyruvic acid-2-C¹⁴, albumin, glycine,
glucose, or sucrose. The av. life of radioactive carotene in
the leaves is 3-5 days. Edwin Howe Scott

Yeznikov, A.A.

VYVAL'KO, I.G.; DUSHECHKIN, A.I. [deceased]; LUSHCHEVSKAYA, G.M.; MATKOVSKIY, K.I.; SAVINOV, B.G.; SHILOV, Ye.A.; YASNIKOV, A.A.

Mechanism of the biosynthesis of isoprene compounds. Part 2. Formation of carotene pigments in the leaves of kok-saghyz. Ukr. khim. zhur. 22 no.5:655-659 '56. (MLRA 10:6)

1. Institut organicheskoy khimii Akademii nauk USSR i Institut fiziologii rasteniy i agrokhimii Akademii nauk USSR.
(Carotene) (Kok-saghyz) (Biosynthesis)

YASNIKOV, A.A.

I-3

USSR/Physiology of Plants. Respiration and Metabolism.

Abs Jour: Ref. Zhur-Biologiya, No 1, 1958, 1142.

Author : Vyval'ko, I.G., Dushechkin, A.I., Matkovskiy, K.I., Shilov, Ye. A., Yasnikov, A.A.

Inst : none given

Title : The Formation in Kok-Sagyz Leaves of Carbonic Acid From Sugars and Acids Containing Marked Carbon

Orig Pub: Byul. po fiziol. rasteniy, 1957, No 1, 15-19.

Abstract: Water solutions of sugars and sodium salts of organic acids containing C^{14} were applied to kok-sagyz leaves (three portions over the course of three days or in one dose before the beginning of the experiment. The air was withdrawn from the plant by use of a vacuum dessicator; the carbon dioxide, absorbable in 25% NaOH, was precipitated, and the radioactivity measured in the dried $BaCO_3$ deposit. It was observed that a considerable

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APPROVED FOR RELEASE: 09/01/2001

CIA-RDP86-00513R001962230001-7"

USSR/Physiology of Plants. Respiration and Metabolism.

Abs Jour: Ref. Zhur-Biologiya, No 1, 1958, 1142.

quantity of $C^{14}O_2$ was isolated two or three hours after introduction of the marked elements. On the fifth or sixth day of the experiment, after application of saccharose C^{14} , α -glucose-1- C^{14} , acetic acid-1- C^{14} , pyruvic acid-2- C^{14} , levulinic acid-1- C^{14} , levulinic acid- C^{14} , and glycolic acid-1- C^{14} in darkness, more than 50% of the introduced radioactivity was isolated. 25% of the radioactivity was isolated upon introduction of -fructose-1- C^{14} and α -gluconic acid 1- C^{14} . In the experiments with levulinic acid C^{14} and pyruvic acid-1- C^{14} , from 33-20% [sic] more $C^{14}O_2$ was isolated in the light than in darkness. As concerns the other substances, the differences between the light and dark variants were imperceptible.

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-6-

YASNIKOV, A. A.

73-1-16/26

AUTHOR: Vyval'ko, I. G., Matkovskiy, K. I. and Yasnikov, A. A.

TITLE: On the Mechanism of Fat Biosynthesis in Maturing Flax Seeds. (K Voprosu O Mekhanizme Biosinteza Zhira v Sozrevayushchikh Semenakh L'na.)

PERIODICAL: Ukrainskiy Khimicheskiy Zhurnal, 1957, Vol. 23, No.1, pp. 85 - 88 (USSR).

ABSTRACT: The distribution of radioactive carbon C^{14} was investigated in fractions of glycerine and fatty acids separated from fats of flax seeds. The transformation of saccharose C^{14} of pyruvic acid-2- C^{14} , acetic acid-1- C^{14} and glyocol-1- C^{14} into fat components, e.g. glycerine and fatty acids of ripening flax seeds was tested by various experiments. The radioactivity of aqueous fractions, of fatty acids and of glycerine is illustrated by graphs 1 - 3. It was shown during the investigations that the acetyl-coenzyme A is synthesised directly from acetic acid and also during the metabolism of carbohydrates and oxidation-reduction processes of acids and aldehydes, e.g. of pyruvic acid. Obviously, the participation of these and other compounds during the synthesis of the acetyl-coenzyme A are linked with the characteristics of the organism or of the living tissues in which these processes take place. E. Newcomb's

Card 1/2.

On the Mechanism of Fat Biosynthesis in Maturing Flax Seeds. 73-1-16/26

et al. conclusions (Ref. 5) are discussed. There are
3 graphs and 7 references, 3 of which are Slavic.

SUBMITTED: May, 3, 1956.

ASSOCIATION: Institute of Plant Physiology and Agrochemistry,
Academy of Sciences, Ukrainian SSR
Institute of Organic Chemistry, Academy of Sciences,
Ukrainian S.S.R. (Institut Fiziologii Rastenii i Agro-
khimii AN USSR. Institut Organicheskoy Khimii AN USSR.)

AVAILABLE: Library of Congress

Card 2/2

YASNIKOV, A.A.

73-2-11/22

AUTHORS: Shilov, Ye.A. and Yasnikov, A.A.

TITLE: The catalytic action of amino acids and amines in organic reactions. 1: Kinetics and enolization of acetone in the presence of amino acids. (Kataliticheskiye deystviya aminokislot i aminov v organicheskikh reaktsiyakh, 1: Kinetika i mekhanizm enolizatsii atsetona v prisutstvi aminokislot).

PERIODICAL: "Ukrainskiy Khimicheskiy Zhurnal" (Ukrainian Journal of Chemistry), Vol.23, No.2, March-April, 1957, pp.215-222 (USSR).

ABSTRACT: The catalytic action of glycine during the enolization of acetone was investigated by A.M.Kuzin and N.A.Nevrayeva (Ref.2:A.M.Kuzin and N.A.Nevraeva:Biokhimiya,1939,Vol.4, 142) in 1939. They measured the velocity of iodination of ketone which is limited by the velocity of its enolization and increases in the presence of acidic and basic agents. (Ref.3:H.Dawson and F.Powis,J.Chem.Soc.,1939,2135; J.Chem.Soc.1926,2282). Kuzin and Nevraeva (2) concluded that the velocity of iodination is proportional to the concentration of glycine and proposed a reaction mechanism for the same. The authors now determined the exact kinetic parameters and iodination (or enolization) of acetone and show that this reaction follows the tri-molecular donor-

Card 1/3

75-2-11/22

The catalytic action of amino acids and amines in organic reactions. 1: Kinetics and enolization of acetone in the presence of amino acids. (Cont.)

acceptor system. Kuzin's and Nevrayeva's mechanism could not be confirmed by these investigations. The kinetics of iodination of acetone in an aqueous solution in the presence of glycine, β -alanine and dimethylglycine were investigated (in the presence or absence of phosphoric acid salts). In all cases, at relatively large and constant concentration of acetone, the reaction proceeds according to the zero order relative to iodine. In the presence of glycine the kinetic equation

$$v = kAGG^-$$

where A = concentration of acetone;
G = glycine in the form of a bipolar ion;
G⁻ = glycocollate ion.

At 25 C: k = 4.0 (units: mole/l min). When both glycine and phosphate are present k' = 410 mole/l min the equation changing to

Card 2/3

$$v = k'AG^- [H_2PO_4^-] + kAGG^- + k_FAF$$

where F = analytic concentration of phosphate salts in

73-2-11/22

The catalytic action of amino acids and amines in organic reactions. 1: Kinetics and enolization of acetone in the presence of amino acids. (Cont.)

moles. For β -alanine $k = 2.74$ and $k' = 604$. Dimethylglycine differs from the 2 above compounds kinetically. The order of the reaction lies between 2 to 1 and the actual effect of phosphate ions becomes negligible. On analyzing the kinetic relations it is found that besides the acetone 2 molecules participate in the transformation of acetone into enol, one having the properties of an acceptor proton and the other of its donor.

There are 2 diagrams, 6 tables and 11 references (4 Slavic).

ASSOCIATION: Institute of Organic Chemistry, Ac.Sc. Ukraine.
(Institut Organicheskoy Khimii AN USSR).

SUBMITTED: November 16, 1956.

AVAILABLE: Library of Congress

Card 3/3

73-2-14/22

73-2-14/22

AUTHORS: Vyval'ko I.G., Matkovskiy K.I. and Yasnikov, A.A.

TITLE: On the mechanism of cellulose biosynthesis in developing flax stems. (K voprosu o mekhanizme biosinteza tsellyulozy v formiruyushchikhsya steblyakh l'na).

PERIODICAL: "Ukrainskiy Khimicheskiy Zhurnal" (Ukrainian Journal of Chemistry), Vol.23, No.2, March-April, 1957, pp.233-235 (USSR).

ABSTRACT: Previously published investigations (1) proved that cellulose is formed in higher plants and some microorganisms at the expense of glucose which is synthesised by these organisms. Experiments carried out by Gritkhaus concerning the study of the biosynthesis of cellulose of cotton fibres. By incising the stem immediately under the boll glucose-1- C^{14} was injected. The thus treated boll ripened in 30 days. Investigations of the distribution of radioactivity showed that 99.97% of the total activity of this sugar occurs in the first C atom. The author concluded that cellulose in cotton fibres is formed by the direct polymerisation of glucose molecules by utilising the energy of phosphorylation. The authors attempted to elucidate the action of saccharose C^{14} , acetic acid-1- C^{14} , glycine-1- C^{14}

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73-2-14/22

On the mechanism of cellulose biosynthesis in developing flax stems. (Cont.)

and pyruvic acid-2-C¹⁴ during the synthesis of cellulose in flax stems. Labelled saccharose-C¹⁴ and glycine-1-C¹⁴ were proved to impart the highest activity. The distribution of radioactivity in cellulose fractions extracted from bast and flax stems is shown in Diagrams 1 and 2. It can be seen that the C¹⁴ atoms pass into the cellulose molecules in various quantities. It is suggested that glycine transforms to d-glucose via the formation of triose.

There are 2 diagrams and 6 references, 3 of which are Slavic.

ASSOCIATION: Institute of Plant Physiology and Agrochemistry, Academy of Sciences, Ukraine and Institute of Organic Chemistry, Academy of Sciences, Ukraine. (Institut Fiziologii Rasteniy i Agrokhimii AN USSR, Institut Organicheskoy Khimii AN USSR).

SUBMITTED: July 12, 1956.

AVAILABLE: Library of Congress

Card 2/2

Yasnikov A.A.

73-3-8/24

AUTHOR: Yasnikov, A. A., Shilov, Ye. A. and Matkovskiy, K. I.

TITLE: Catalytic Action of Amino Acids and Amines in Organic Reactions. 2. Kinetics and Mechanism of Enolisation of Levulinic Acid in the Presence of Glycine and Other Amino Acids. (Kataliticheskiye Deystviya Aminokislót i Aminov v Organicheskikh Reaktsiyakh. 2. Kinetika i Mekhanizm Yenolizatsii Levulinovoy Kisloty v Prisutstvii Glikokola i Drugikh Aminokislót)

PERIODICAL: Ukrainskiy Khimicheskii Zhurnal, 1957, Vol. 23, No. 3, pp. 333-340 (USSR).

ABSTRACT: The first communication gave results of investigations on the kinetics of enolisation of acetone in the presence of amino acids. (Ref. 1) Data on the kinetics of iodination of levulinic acid under the above mentioned conditions are investigated in the present article. Levulinic acid differs from acetone in the fact that it contains an acid group which occurs in 2 forms: the undissociated acid (the concentration being denoted as M) and the acid ion (M^-). M = total concentration of the undissociated acid and its ion. Investigations have shown that the iodination of levulinic acid in the presence of amino acids is limited by the enol formation stage. It can be assumed that 3- or 5-iodolevulinic acid is formed in the first instance. The structure of the iodination products

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73-3-8/24

Catalytic Action of Amino Acids and Amines in Organic Reactions.
2. Kinetics and Mechanism of Enolisation of Levulinic Acid in the
Presence of Glycine and Other Amino Acids.

was not investigated. The method of measuring the rate of iodination was described in part 1. (Ref. 1). Direct titration in acidified solutions was used. This method of titration was found suitable. The synthesised levulinic acid was subjected to 2 vacuum distillations. Fractions boiling at 137 - 139°C at 10 mm Hg. Iodine sublimes from the pulverised mixture with potassium iodide. Chemically pure glycine was used for the experiment. Figure 1 shows that the rate of iodination of levulinic acid increases with increasing concentration of glycine and of the hydroxyl-ion. The rate of iodination does not depend on the concentration of the iodine, G_{an} and is, as in the case of acetone, defined by the stage of enolisation according to the kinetic equation: $v = k_0 M$. At comparatively high concentrations of levulinic acid the reaction rate is represented by straight lines. At 0.5 mole concentrations of levulinic acid, in the presence of 0.005 mole I_2 and 1.25 mole glycine a fast reaction occurs until 10% of I_2 is spent (Fig. 2).

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73-3-8/24

Catalytic Action of Amino Acids and Amines in Organic Reactions.
2. Kinetics and Mechanism of Enolisation of Levulinic Acid in the Presence of Glycine and Other Amino Acids.

Thereafter the reaction rate decreases. In the presence of glycine the rate of enolisation of levulinic acid depends on the pH of the solution (viz. Fig. 3). At low pH (4.5 - 6) the constant k_0 does not increase appreciably with increasing pH, with more alkaline solutions the rate increases sharply. Table 1 gives values for k_0 and $k_2'K_2$ (where $K_2 = 2.5 \times 10^{-10}$). The addition of buffer salts (diphthalates) denoted as B in the table, does not affect appreciably the aforementioned values. The value of $k_2'K_2$ at pH 4.8 - 6 equals 1.1×10^{-8} , i.e. $k_2' = 44$. Just as in the calculation of acetone the authors found that the concentration of glycine appears in the equation as G^2 . Table 2 gives the calculated values for kK_2 where

$$k = \frac{k_0 a_{N+}}{K_2 G^2}$$

At pH 6 the rate of enolisation of levulinic acid increases with the concentration of the added buffer phosphate salt (F). Figure 1 and Table 3 show the relation between k_0 and pH in the presence of phosphates. Several experiments

Card 3/4

73-3-8/24

Catalytic Action of Amino Acids and Amines in Organic Reactions.
2. Kinetics and Mechanism of Enolisation of Levulinic Acid in the
Presence of Glycine and Other Amino Acids.

with various amino acids in the presence of phosphates at 25°C were carried out in order to compare their catalytic action. Table 4 gives the k_0 values for these experiments. Table 5 gives data for the relation of the rate of enolisation and the temperature. The energy of activation was 18750 cal. The kinetic equation for pH 4.5 - 9 was found to be

$$v = k_2' M_K G^- + k M G G^- + k' M G^-(H_2PO_4^-).$$

At 25°C: $k_2' = 44$; $k = 0.68$; $k' = 64$ units: mole/litre, minute). The donor - acceptor scheme of enolisation of ketones in the presence of amino acids and proton donors is discussed. There are 3 figures, 5 tables and 5 references, 2 of which are Slavic.

SUBMITTED: November, 18, 1956.

ASSOCIATION: Institute of Organic Chemistry, Academy of Sciences
Ukrainian SSR. (Institut Organicheskoy Khimii AN USSR)

AVAILABLE: Library of Congress.
Card 4/4

Yasnikov A.A.

VYVAL'KO, I.G.; DUSHECHKIN, A.I. [deceased]; LUSHCHEVSKAYA, G.M.;
MATKOVSKIY, K.I.; SAVINOV, B.G.; SHILOV, Ye.A.; YASNIKOV, A.A.

Mechanism of the biosynthesis of isoprene compounds. Part 3:
New data on rubber and carotene antecedents in kok-saghyz.
Ukr.khim.zhur. 23 no.4:516-522 '57. (MIRA 10:10)

1. Institut organicheskoy khimii AN USSR i Intitut zemledeliya
Ukrainskoy akademii sel'skokhozaystvennykh nauk.
(Biochemistry) (Kok-saghyz)

YASNIKOV, A.

The following papers were presented at the Annual Meeting of the German Chemical Society in Leipzig, 28 October - 1 November 1958:

BALANDIN, A. A. (AS USSR, Moscow)

"Toward the Construction of a Unified Theory of Catalysis - Structure and Energy Factors."

BORESKOV, G. K. (Physico-Chemical Inst. in Karpov, Moscow)

"Reciprocal Influence of Reagent Systems and Catalysts."

BRAUNSHTEIN, A. Ye. (Inst. of Biological and Medical Chemistry, AMU, USSR, Moscow)

"General Questions of Biological Catalysis in the Light of the Activity of the Pyridoxalenzym."

EMANUEL, E. (Inst. Chemical Physics, AS USSR, Moscow)

"Homogeneous Catalysis and Chemical Induction in Slow Oxidation - Chain Reaction."

PODINSKIY, S. Z.

"Mechanism and Kinetics of the Half-Life Catalyst."

SHILOV, E. and YASNIKOV, A. (Inst. Organic Chemistry, AS Ukrainian SSR)

"Mechanism of Catalytic Action of Aminocids in the Reaction of Carbonyl Compounds."

VILESOV, P. and TERNIN, A. (Physics Inst., Leningrad State University)

"Photoelectric Emission on the Surface of Half-Life Catalysts."

SO: Chemische Technik, October 1958, Ussr.

VYVALKO, I.O.; MATKOVSKIY, K.I.; YASHNIKOV, A.A.

Oil biosynthesis in the maturing linseed. Ukr. khim. zhur. 24
no.3:372-374 '58. (MIRA 11:9)

1. Institut semledeliya Ukrainskoy Akademii sel'skokhozyaystvennykh
nauk i Institut organicheskoy khimii AN USSR.
(Linseed oil) (Biosynthesis)

VYVAL'KO, I.G.; DUSHCHKIN, A.I. [deceased]; LUSHCHENSKAYA, G.M.; MATKOVSKIY, K.I.;
SAVINOV, B.G.; SHILOV, Ye.A.; YASHNIKOV, A.A.

Biosynthesis of carotene. Vitaminy no.4:159-163 '59.
(MIRA 12:9)

1. Institut organicheskoy khimii Akademii nauk USSR i Institut
zemledliya Ministerstva sel'skogo khozyaystva USSR, Kiev.
(CAROTENE)

YASNIKOV, A.; SHILOV, E.

Structure of the stibromophenol bromide. In Russian. p. 63.

ACTA CHIMICA. (Magyar Tudományos Akadémia) Budapest. Hungary
Vol. 21, no. 1. 1959.

Monthly list of East European Accessions (EEAI) LC Vol. 9, no. 2, Feb. 1960
encl.

17(3)

AUTHORS:

Shilov, Ye. A., Academician of the
AS UkrSSR, Yasnikov, A. A.

SOV/20-124-2-61/71

TITLE:

On the Part Alanine Plays in Biosynthetic Processes in Plants
(Ob uchastii alanina v biosinteticheskikh protsessakh v
rasteniyakh)

PERIODICAL:

Doklady Akademii nauk SSSR, 1959, Vol 124, Nr 2,
pp 459 - 461 (USSR)

ABSTRACT:

Alanine belongs to the substances forming during the initial stages of photosynthesis in plants (Ref 1). Little is known about its further transformations, especially about the problem mentioned in the title. It is a remarkable fact that in the liver and the kidneys alanine may be a better donor of acetyl groups than acetic-, lactic- and pyruvic acid (Refs 3, 4). Alanine is of greatest importance as an intermediate product in the formation of rubber and carotene in Kok-Sagyz (a rubber producing type of composite flowers from Central Asia) (Ref 5), as well as of fats in linseed (Ref 6).

Card 1/3

Experiments proved that the alanine-2-C¹⁴ or -3-C¹⁴ introduced

On the Part Alanine Plays in Biosynthetic Processes
in Plants

SOV/20-124-2-61/71

through leaves or roots can provide rubber with a higher radioactivity than it is the case with radioactive acetic acid or saccharides etc, under the same conditions. This holds also for carotene which, however, becomes also radioactive in the case of introduction of marked acetic acid. If, however, carbon in the alanine molecule in the carboxyl group is marked neither rubber nor carotene become remarkably radioactive. Figures 1 and 2 give as an example results of an experimental series. In other experimental series (Ref 5) alanine-2- 14 C caused in rubber a 3 to 4 times higher activity than acetic acid, glycol, acetoacetic acid and several hexoses. The l- and d-types of alanine take part to about the same extent in the formation of radioactive rubber as their racemic mixture. The results obtained prove that the 2-carbon residue forming as a result of the decarboxylation of alanine plays a part in different biosynthetic processes in plants. Alanine is obviously transformed into carotene molecules passing through the stage of pyruvic- and acetic acid (Fig 2). The higher rapidity and the degree of absorption of the alanine-2-carbon residue into rubber compared to acetic acid

Card 2/3

On the Part Alanine Plays in Biosynthetic Processes
in Plants

SOV/20-124-2-61/71

point to the fact that the forming alanine decarboxylation product (apparently acetaldehyde imine) is not transformed into an acetyl coenzyme A but that it takes immediately part in the formation of the molecules of rubber or its predecessors. The participation of alanine in several biosynthetic processes in plants is obviously due to the fact that alanine can serve as donor of 2-carbon acids: acetic acid and acetaldehyde imine. The latter form further links of the chain of biosynthetic reactions. There are 3 figures and 11 references, 4 of which are Soviet.

ASSOCIATION: Institut organicheskoy khimii Akademii nauk USSR (Institute of Organic Chemistry of the Academy of Sciences, UkrSSR)

SUBMITTED: September 22, 1958

Card 3/3

SHILOV, Ye. A.; VAYNSHTEYN, F.M.; YASNIKOV, A.A.

Participation of halogen cations in the mechanism of halogenation
of organic compounds. Kin. i kat. 2 no.2:214-220 Kr-Ap '61.
(MIRA 14:6)

1. Institut organicheskoy khimii AN USSR, Kiev.
(Halogenation)

YASNIKOV, A.A.; GAYVORONSKAYA, Ye.M.

Catalytic action of amino acids and amines in organic reactions.
Part 3: Kinetics and mechanism of the formation and decomposition
of benzalmalonic esters in the presence of amino acids. Ukr. khim.
zhur. 27 no.4:506-518 '61. (MIRA 14:7)

1. Institut organicheskoy khimii AN USSR.
(Malonic acid) (Amino acids)

SHILOV, Ye.A.; YASNIKOV, A.A.

Catalytic action of amino acids and amines in organic reactions.
Part 4: Specific catalytic action of amines in the enolization
of ketones. Ukr.khim.zhur. 27 no.5:639-658 '61. (MIRA 14:9)

1. Institut organicheskoy khimii AN USSR.
(Amines) (Catalysis) (Ketones)

YASNIKOV, A.A.

Kinetics and mechanism of the hydrolysis of ethenolpyruvic
acid phosphate. Ukr.khim.zhur. no.3:343-357 '61.
(MIRA 14:11)

1. Institut organicheskoy khimii AN USSR.
(Pyruvic acid)
(Phosphoric acid)
(Hydrolysis)

YASINOV, A.A.; BOYKO, T.S.

Catalytic action of amino acids and amines in organic reactions.
Part 6. Kinetics and mechanism of the formation and decomposition
of diethyl ester of butylidenemalononic acid in the presence of
amino acids. Ukr.khim.zhur. 27 no.6:777-782 '61.
(RITA 14:11)

1. Institut organicheskoy khimii AN USSR.
(Malonic acid)
(Amino acids)

YASNIKOV, A.A.; MATKOVSKIY, K.I.; GAYVORONSKAYA, Ye.M.

Catalytic action of amines and amino acids. Part 4. Kinetics and mechanism of crotonic condensation of butyraldehyde in the presence of glycocoll. Ukr. khim. zhur. 28 no.1:88-94 '62. (MIRA 16:8)

1. Institut organicheskoy khimii AN UkrSSR.

YASNIKOV, A.A.; MATKOVSKIY, K.I.

Catalytic action of amino acids and amines in reactions of carbonyl compounds. Part 7: Condensation of the enamine base of N-propenylpiperidine with propionaldehyde and acetone. Ukr.khim. zhur. 28 no.2:210-212 '62. (MIRA 15:3)

1. Institut organicheskoy khimii AN USSR.
(Piperidine) (Propionaldehyde) (Acetone)

YASNIKOV, A. A.; BOYKO, T. S.

Catalytic action of amino acids and amines in organic reactions.
Part 8: Kinetics and mechanism of acetylation of glycocoll by
S-acetylthioglycolic acid, Ukr. khim. zhur. 28 no.3:347-358
'62. (MIRA 15:10)

1. Institut organicheskoy khimii AN UkrSSR.

(Glycine) (Acetylation) (Catalysis)

YASNIKOV, A. A.; BOYKO, T. S.

Catalytic action of amino acids and amines in organic reactions.
Part 9: Acetylation of aniline by acetylthioglycolic acid. Ukr.
khim. zhur. 28 no.5:624-626 '62. (MIRA 15:10)

1. Institut organicheskoy khimii AN Ukr-SSR.

(Aniline) (Acylation)

GRISHIN, O. N.; YASNIKOV, A. A.

Kinetics and mechanism of the addition of weak acid anions to N-benzyl-1,4-dihydronicotinamide. Ukr. khim. zhur. 28 no.6:707-713
'62. (MIRA 15:10)

1. Institut organicheskoy khimii AN UkrSSR.

(Acids) (Nicotinamide)

BOYKO, T.S.; VOLKOVA, N.V.; YASNIKOV, A.A.

Catalytic action of amino acids and amines in organic reactions.
Part 10: Specific action of amino acids and amines in crotonic
and aldol condensations of butyraldehyde. Ukr. khim. zhur. 29
no.11:1179-1187 '63. (MIRA 16:12)

1. Institut organicheskoy khimii AN UkrSSR.

VOLKOVA, N.Y.; YASNIKOV, A.A.

Mechanism of the condensation of enamines with aldehydes.
Dokl.AN SSSR 149 no.1:94-96 Mr '63. (MIRA 16:2)

1. Institut organicheskoy khimii AN UkrSSR. Predstavleno
akademikom B.A.Kazanskim.
(Amines) (Aldehydes)

MEL'NICHENKO, I.V.; SHILOV, Ye.A.; YASNIKOV, A.A.

Catalytic action of amino acids and amines. Part II: Kinetics of enolization of acetone in the presence of amino acids of various structure. Ukr. khim. zhur. 30 no.6:599-608 '64. (MIRA 18:5)

1. Institut organicheskoy khimii AN UkrSSR.

MEL'NICHENKO, I.V.; YASNIKOV, A.A.

Catalytic action of amino acids and amines in organic reactions.
Part 12: Catalytic action of imidazole in acetone enolization.
Ukr.khim. zhur. 30 no.7:723-728 '64 (MIRA 18:1)

1. Institut organicheskoy khimii AN UkrSSR.

MEL'NICHENKO, I.V.; YASNIKOV, A.A.

Catalytic action of amino acids and amines. Part 13: Specific
action of histidine and histamine in aldol condensation of
butyraldehyde. Ukr. khim. zhur. 30 no.8:838-847 '64. (MIRA 17:11)

1. Institut organicheskoy khimii AN UkrSSR.

MEL'NICHENKO, I.V.; SHILOV, Ye.A.; YASNIKOV, A.A.

Catalytic effect of amino acids and amines in organic reactions.
Part 14: Mechanism of hydration of crotonic aldehyde. Ukr.khim.
zhur. 30 no.11:1171-1178 '64. (MIRA 18:2)

VOLKOVA, N.V.; YASNIKOV, A.A.

Kinetics and mechanism of hydrolysis of acetol phosphate. Ukr.
khim.zhur. 30 no.11:1178-1183 '64. (MIRA 18:2)

1. Institut organicheskoy khimii AN UkrSSR.

YASHNIKOV, A.A.; SHILOV, Ye.A.; VOLKOVA, N.V.

Catalytic action of amino acids and amines in organic reactions.

Part 15. Ukr. khim. zhur. 30 no.12:1313-1318 '64
(MIRA 18:2)

1. Institut organicheskoy khimii AN UkrSSR.

VOLKOVA, N.V.; YASNIKOV, A.A. [IAsnikov, O.O.]

Kinetics and mechanism of acetol phosphate iodination in the
presence of amines and amino acids. Dop. AN URSR no.7:901-903
'65. (MIRA 18:8)

1. Institut organicheskoy khimii AN UkrSSR.

VOLKOVA, N.V.; SHILOV, Ye.A.; YASNIKOV, A.A.

Catalytic action of amino acids and amines in reactions of carbonyl compounds. Part 16: Kinetics and mechanism of iodination of acetol phosphate in the presence of ethylenediamine. Ukr. khim. zhur. 31 no.1:56-66 '65. (MIRA 18:5)

1. Institut organicheskoy khimii AN UkrSSR.

YASNIKOV, A.A.; SHILOV, Ye.A.; GRISHIN, O.M.

Reduction of benzaldehyde and p-nitrobenzaldehyde by N-benzyl-1,
4-dihydronicotinamide in an alkaline solution. Ukr. khim. zhur.
31 no.1:118-119 '65. (MIRA 18:5)

1. Institut organicheskoy khimii AN UkrSSR.

VOLKOVA, N.V.; YASNIKOV, A.A.

Products of phosphorylation of acetol by phosphorus oxychloride
in quinoline. Ukr. khim. zhur. 31 no.1:119-120 '65. (MIRA 18:5)

1. Institut organicheskoy khimii AN UkrSSR.

VOLKOVA, N.V.; MEL'NICHENKO, I.V.; YASNIKOV, A.A.

Catalytic action of amino acids and oximes in reactions of
carbonyl compounds. Part 17: Catalytic action of lysine in
iodination of acetol phosphate and acetone. Ukr. khim. zhur.
31 no.9:936-942 '65. (MIRA 18:11)

1. Institut organicheskoy khimii AN UkrSSR.

AUTHOR: ~~Yasnikov, L.Ye.~~

SOV/132-58-12-9/14

TITLE: The Correction "For the Arrester" in Polyakov's Device (Popravka "za arretir" k priboru Polyakova)

PERIODICAL: Razvedka i okhrana nedr, 1958,²⁴ Nr 12, pp 48-49 (USSR)

ABSTRACT: When Polyakov's device is used without making the necessary adjustments during drilling operations, the bore holes can deviate as much as 20 - 25 m. This deviation can distort the evaluation of the deposit. The author proposes to prepare a table of corrections for every device individually and to have them periodically checked. There are 2 diagrams and 1 table.

ASSOCIATION: Ulenskaya geologorazvedochnaya partiya (The Ulen' Geological Prospecting Party)

Card 1/1

YASNIKOY, N. G.

YASNIKOY, N.G., kandidat sel'skokhozyaystvennykh nauk.

Sowing grain in hills on irrigated land. Zemledelia 5 no.8:72-75
Ag '57. (MIRA 10:9)

(Sowing) (Grain)

YASNIKOV, V.N.

Dec 1947

USSR/Electricity
Power lines
Circuit Breakers

"Operation of Power Lines with Phase Control and Single-Pole Automatic Repeater-Type Break-In Systems, " G.I. Lysakovskiy, L. Ye. Nebrat, V.N. Yasnikov, Engineers, $\frac{1}{2}$ pp

"Elektricheskiye Stantsii" No 12

Recently, second power line of the KemerovoEnergo system changed over the phase control, 110-kw line running 83.6 km. Authors briefly describe other aspects of the circuit, such as automatic repeater-type break-in systems, and relays connected to this phased power line,

PA 50T11

YASHNIKOV, V. N.

PA 53/49T32

USSR/Electricity
Transmission Lines
Electric Power Systems

Jun 49

"Exploitation of 115-Kilovolt Lines Equipped
With Separate Phase Control and OAPV (Separate
Automatic Repeated Switching)," I. Ye. Nebret,
V. N. Yashnikov, Engineers, 3 pp

"Elek Stants" No 6

Sum up the work of a power system using a num-
ber of 115-kv transmission lines equipped with
separate phase control and OAPV 80 - 100 km
long. In 5 years of exploitation, there were
about 30 cases of automatic switching out of

53/49T32

USSR/Electricity

(Contd)

Jun 49

separate phases during frosts, breaks in con-
ductors, damage to supports, etc. In all cases
except three, power supply to consumers was
maintained on two phases.

53/49T32

PA156T24

YASNIKOV, V.N.

USSR/Engineering - Power Plants, Electric
Electrical Equipment

Nov 49

"Some Design Requirements for Siberian Rayon Electric
Power Stations," L. Ye. Nebrat, V. N. Yasnikov, Engi-
neers, Stalin Prize Laureates, 3 pp

"Elek Stants" No 11

Siberian electric power stations are designed, as r.
rule, on same principles as those in other parts of
USSR. Practice shows, however, that special atten-
tion must be paid to severe climatic conditions.
Lists 19 points in connection with fuel supply, water
supply, powerhouse, electrical equipment, and oil sup-
ply.

156T24

RUDEKNO, Yu.N., kand.tekhn.nauk (Kemerovo); YASNIKOV, V.N., inzh. (Kemerovo)

Determination of conditions for the optimum operation of electric
power systems. Elektrichestvo no.4:90-91 Ap '62. (MIRA 15:5)
(Electric power distribution)

PORTNOY, M.G., kand. tekhn. nauk; RUDENKO, Yu.N., kand. tekhn. nauk;
YASNIKOV, V.N., inzh.

Joining of Siberian power distribution networks by weak intersystem
couplings. Elek. sta. 34 no.5:37-43 My '63. (MIRA 16:7)

(Siberia--Interconnected electric utility systems)

LYTAYEV, I.A., inzh.; RUDENKO, Yu.N., kand. tekhn. nauk; YASHNIKOV, V.N., inzh.

Features of using automatic frequency control in unified
power systems. Elek. sta. 36 no.9:26-30 S '65. (MIRA 18:9)

RUDENKO, Yu.N., kand. tekhn. nauk; SOKOLOV, V.K., inzh.; YASNIKOV,
V.N., inzh.

Distribution of power reserves in an electric power system.
Elek sta. 35 no.10:43-49 0'64. (MIRA 17:12)

PRIVALOVA, S.G., inzh.; ROBINOV, Yu.M., inzh. tekhn. nauk: 1955. ...
inzh.; YASHINOV, V.H., inzh.

Statistical forecast of the load graph of electric power systems.
(KIB 1955)
Elek. stn. 36 no.1:62-65 Jan '65.

TOLMACHOV, I.M.; BOGDAN, I.K. [Bohdan, I.K.]; YASNIKOVA, O.O.

Respiration of wintering plants. Nauk. pratsi UASHN 17 no.12:
30-33 '60. (MIRA 16:7)

(Plants—Respiration)
(Plants. Effect of cold on)

L 42878-66 EWT(m)/T/EWP(t)/ETI IJP(c) JD/WW/JH/JWD

ACC NR: AP6022894

SOURCE CODE: UR/0078/66/011/004/0886/0889

AUTHOR: Berg, L. G.; Yasnikova, T. Ye.

ORG: Kazan State University im. V. I. Ul'yanov-Lenin (Kazanskiy gosudarstvennyy universitet)

TITLE: Thermographic determination of the heats of polymorphic transformations

SOURCE: Zhurnal neorganicheskoy khimii, v. 11, no. 4, 1966, 886-889

TOPIC TAGS: phase transition, heat of transition, thermogram

ABSTRACT: An experimental study was made to determine the heats of polymorphic transformations recorded on various thermograms. The calculation was carried out by using the formula

$$\frac{Q_1}{Q_2} = x \frac{S_1}{S_2}$$

where Q_1 and Q_2 are the heats of the phase transformations, S_1 and S_2 are the reduced areas of differential peaks, and x is a conversion factor. The chosen substance with a known heat of polymorphic transformation was KNO_3 . The substances studied were NH_4Br , NH_4Cl , $AgNO_3$, and NH_4NO_3 (which had three different heats of transformation), and the corresponding x values, determined directly from the thermograms, were found

Card 1/2

UDC: 536.65:541.7

L 42878-66

ACC NR: AP6022894

to be in good agreement with data reported in the literature. Orig. art. has: 2 figures, 1 table, and 4 formulas. 0

SUB CODE: 0720/SUBM DATE: 16Jan65/ ORIG REF: 005/ OTH REF: 005

Card 2/2 *klh*

YASNIKOVA, YE. A.

"Natural Serinal Propagation of Main Foliate Species Beneath the Canopy of Planted Vegetation in the Steppes of the Southeast." Acad. Sci. Belorussian SSR, Inst. of Socialist Agriculture, (Minsk, 1955). (Dissertation for the Degree of Candidate in Agricultural Sciences)

SO: Knizhnaya Letopis', No. 22, 1955, pp 93-105

PIVOVAROV, S.P.; FIRSOV, Ye.P.; YASNILO, O.N.; LATYSHEV, G.D.

Comparison of circuits for paramagnetic resonance detection.
Trudy Inst. iad. fiz. AN Kazakh. SSR 6:119-123 '63.
(MIRA 16:10)

YASNITS, Mikhail Grigoryevich

PHASE I BOOK EXPLOITATION

SOV/1265

11(0)

Kamyshov, Sevast'yan Filippovich, Galikhin, Viktor Dmitriyevich, Larin
Vasily Il'ich, Mikhaylov, Leonid Leonidovich, Filonova, Lidiya Ivanovna,
Yasnits, Mikhail Grigor'yevich, and Kvochkin, Fedor Abramovich

Groznenskaya neftyanaya promyshlennost' (The Grozny Petroleum Industry) Moscow,
Gostoptekhizdat, 1957. 57 p. 1,500 copies printed.

Executive Ed.: Lozbyakova, Ye. S.; Tech. Ed.: Polosina, A.S.

PURPOSE: The book is intended for engineers, technicians and workers in the
petroleum industry.

COVERAGE: The status of the Grozny petroleum industry before the Revolution and
the achievements in the recovery and refining of petroleum during the 40 years
after the Revolution are discussed. New oil fields, petroleum installations
and modern techniques and procedures introduced in the Grozny petroleum indus-
try are described. No facilities are mentioned. No references are given.

Card 1/3

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11(0)

The Groznyy Petroleum Industry

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11(0)

SOV/1265

The Groznyy Petroleum Industry

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AVAILABLE: Library of Congress

Card 3/3

TM/mas
3-19-59

L 9896-66 EWT(1)/EWA(d)/T-2/EWA(m)-2 IJP(c) AT

ACC NR: AP5026573

SOURCE CODE: UR/0281/65/000/005/0091/0104

AUTHOR: Tolmach, I. M. (Khar'kov); ^{44.55}Yasnitskaya, N. N. (Khar'kov)

ORG: none

TITLE: Hall effect in a channel with section electrodes

SOURCE: AN SSSR. Izvestiya. Energetika i transport, no. 5, 1965, 91-104

TOPIC TAGS: magnetohydrodynamics, Hall effect

ABSTRACT: The current distribution in a ^{21, 44, 55}magnetohydrodynamic (MHD) generator having section electrodes and an arbitrary value of the Hall parameter is ^{16.71.5}described analytically. Examination of a nonperiodic problem (Gilbert's boundary problem for a strip) results in the Keldysh-Sedov formula for current distribution. When the number of electrodes is very great, the Gilbert problem becomes a boundary problem for dual-periodic functions. Formulas are derived for the

Card 1/2

UDC: 621.3.013:532.501.33

2

L 9896-66

ACC NR: AP5026573

arbitrary constants, in the current-distribution relations, which depend on the connections of electrodes. Curves of the MHD generator efficiency, output, and voltage vs. the flow-width-to-section-pitch ratio are plotted. Two extensive supplements present details of the mathematical operations involved. Orig. art. has: 5 figures and 107 formulas.

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PC
Card 2/2

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[Mineral waters of the southern part of Eastern Siberia] Mineral'nye vody iuzhnoi chasti Vostochnoi Sibiri. Moskva. Vol.1. [Hydrogeology of mineral waters and their significance for the national economy] Gidrogeologiya mineral'nykh vod i ikh narodnokhoziaistvennoe znachenie. Pod obshchei red. V.G.Tkachuk i N.I.Tolstikhina. 1961. 346 p. (MIRA 14:8)

1. Akademiya nauk SSSR. Sibirskoye otdeleniye. Vostochno-sibirskiy geologicheskii institut. (Continued on next card)

TKACHUK, V.G.--- (continued) Card 2.

2. Vostochno-Sibirskiy geologicheskii institut (for Tkachuk, Firneker, Yasnitskaya, Krutikova, Lysak). 3. Institut geografii Sibirskogo ot-deleniya Akademii nauk SSSR (for Shotskiy). 4. Chitinskoye geologiches-koye upravleniye (for Orlova). 5. Sosnovskaya ekspeditsiya Mini-sterstva geologii i okhrany neдр SSSR (for Stepanov). 6. Irkutskiy gosudarstvennyy universitet (for Vlasov, Prokop'yev, Chernyshev, Pav-lova). 7. Leningradskiy gornyy institut (Tolstikhin). 8. Gosudar-stvennyy nauchno-issledovatel'skiy institut kurortologii i fiziote-rapii (for Ivanov, Yarotskiy, Karaseva, Arutyunants, Romanova). 9. Irkutskoye geologicheskoye upravleniye (for Trofimuk). 10. Bay-kal'skaya limnologicheskaya stantsiya Vostochno-Sibirskogo filiala AN SSSR (for Ladeyshchikov). 11. Otdel ekonomiki i geografii Vostochno-Sibirskogo filiala AN SSSR (for Kruchinina).

(Siberia, Eastern--Mineral waters)

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I.S., kand. geol.-miner. nauk, red.; PINNEKER, Ye.V.,
kand. geol.-miner. nauk, red.; YASHITSKAYA, N.V., red.;
FILIPPOVA, B.S., red.; SHOKHET, B.S., red.izd-va;
GUS'KOVA, O.M., tekhn. red.

[Mineral waters of Eastern Siberia] Mineral'nye vody
Vostochnoi Sibiri. Moskva, Izd-vo AN SSSR, 1963. 148 p.
(MIRA 17:1)

1. Akademiya nauk SSSR. Sibirskoye otdeleniye. Institut
zemnoy kory.

TURKIN, P.P.; YASNITSKAYA, T.A.

True adhesion of the placenta. Akush.i gin. no.2:72 Mr-Ap '53. (MLRA 6:5)
(Placenta--Diseases)

Chemistry and technology of sulfonamide preparations.

1. Causes of unsatisfactory filtration of sulfonamide chloride precipitates and methods of its acceleration. II. Vn. Yanitskii. *J. Applied Chem. (U.S.S.R.)* 10, 1313-18 (1946) (in Russian).—The sulfonamide chlorides as intermediates for the production of sulfonamides, usually obtained from acylated amines, are readily hydrolyzed in the presence of H₂O, acid, and high temp.; rapid filtration is imperative, but these chlorides filter very poorly. Diphenylurea disulfonamide chloride is particularly bad in this respect. Observation of a no. of ppt. in aq. of various electrolytes led to the conclusion that the sulfonamide chlorides form a coarsely dispersed hydrophilic suspension carrying a neg. charge. It was apparent that the elec. charge played a great role in the stability of the suspension. The use of different electrolytes made it apparent that the product obeys the law of Schulze-Hardy (the discharge efficiency of ions carrying an opposite sign is dependent on its valence).

Addn. of NaCl, CaCl₂, AlCl₃, FeCl₃, and PbNH₄ in 0.1 M concn. gave a filtration time of 7, 4, 2, 1.5, and 4 min., resp. (all ppts. were sucked dry for 30 min.), giving a very dense cake, dry to the touch, contg. 55.4, 38.2, 36.8, 37.6, and 55.0% moisture; the control required 100 min. to filter, and had a moisture content of 72.5%; the same effect was achieved by passing a steady current of 0.00 v. through the suspension. Increasing the concn. of AlCl₃ to 2 M increased the filtration time to 25 min. (from 1.5) and the moisture to 46.3%, apparently due to a recharge of the particles by the polyvalent ion. The most desirable were the trivalent metals Al and Fe; the concn. was small

enough not to cause contamination. Temp. played an important role; filtration at 0°, 20°, 30°, and 40° lasted 110, 90, 40, and 2 min., resp., and the cakes had a moisture content of 78.6, 68, 46.3, and 37.8%, resp. The stability of the susp. at 40° was good in acid medium, no decomp. occurring between 1 and 24 hrs.; the ppt. was a coarse aggregate, with less surface and greater stability toward hydrolysis. Org. cations are readily adsorbed by the org. micelle, forming a hydrophobic surface resulting in more rapid aggregation. A more easily hydrolyzed acid amide liberating the corresponding amine, filtered much faster. According to the law of periodicity (Petchenko-Kel'tchenko, *J. Gen. Chem.* 3, 59 (1933)), a mol. of urea contg. 2 identical substituents is less subject to hydrolysis than acetamide and formamide. The sulfonamide chlorides of diphenylurea (one questions the hydrolysis of this amide linkage before the Cl), acetanilide, phenylurethane, and formamide were found to hydrolyze 2.14, 14.4, 16, and 18%, resp., the filtration times of the products being 90, 2, 2.5, and 2.3 min., and the moisture contents 68, 28.2, 27.8, and 40% (the degree of hydrolysis of the acyl group, expressed in % of the starting amine, was detd. by diazotization of the filtrate); but the hydrolysis of sulfonamide chlorides leading to the formation of org. anions (AcNH₂Cl₂SO₂⁻) will peptize the ppt., forming a hydrophilic surface and particles difficult to filter. Boris Gutof

ca

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Diphenylureadisulfonamide. H. Ya. Yasnitskiy, U.S.S.R., 60,675, July 31, 1940. $\text{C}_{12}\text{H}_{10}\text{N}_2\text{O}_4\text{S}_2$, is treated with aq. NH_4OH in the presence of electrolytes such as NaCl , NH_4Cl , or $(\text{NH}_4)_2\text{SO}_4$. The disulfamide thus obtained is large-grained, filters readily, and is easy to wash. The yield is 90-92% of theoretical.

M. Howch

ASAC-SLA METALLURGICAL LITERATURE CLASSIFICATION

SERIES	ALPHA	BETA	GAMMA	DELTA	EPSILON	ZETA	ETA	THETA	IOTA	KAPPA	LAMBDA	MU	NU	XI	OMICRON	PICHA	RHO	SIGMA	TAU	UPSILON	PHI	CHI	PSY	OMEGA
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	

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G
Chloral. R. I. Vall and B. Ya. Yasnitskii. U.S.S.R.
No. 720, Nov. 30, 1947. Chloral is produced by catalytic
chlorination of EtOH. As catalyst is used Cl compds.
of S, P, or I. M. Hosh

BA

6

Some comments on paper by L. S. Solodov and E. N. Shvachkin:
 "Rate of synthesis and in the process of chlorosulphonation of
 acetanilide." B. Zh. Khim. (U.S.S.R., 1950, 26, 853-854). A series of mistakes (now are mentioned) and
 contradicting data in Solodov and Shvachkin's paper (B., 1950,
 26, 853-854) cause doubts as to the correctness of some of the conclusions.
 It is particularly misleading to assume that the role of H_2SO_4 in
 the chlorosulphonation of acetanilide is of "secondary" importance.
 The role of H_2SO_4 is very substantial because it decreases the
 concn. of $Cl-SO_3H$ and increases the final yield of sulphonide
 by shifting to the right the equilibrium of the reaction: p -
 $NHAc-C_6H_4-SO_3Cl + H_2SO_4 \rightarrow p-NHAc-C_6H_4-SO_3H + HCl + SO_2$
 J. B. J. ZABA.

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Rate of sulfuric acid in the chlorosulfonation of *o*-alkyl-
anilides. B. Ya. Yanitskii (Kharkov Inst. Pharm. Chem.
Research). *J. Applied Chem. U.S.S.R.* 23, 945-8 (1950)
(Engl. translation); cf. Sokolov and Shevchenko, *C.A.* 44,
2406v. — Four objections are raised concerning the assump-
tions made by S. and S. in the study of the reaction of
PhNHAc (I) with H₂SO₄ (II) and ClSO₃H (III): 1. I and
II give H₂O, which decomps. III and changes the concn. of
both II and III; the kinetics of reaction of I with II and I
with III cannot be solved on the basis of a simple equation.

The problem is further involved in that HSO₃H (IV) reacts
with III, and HSO₃Cl (V) reacts with II. 2. Only the
ratio of the concns. in the final state det. the yields of IV and
V, since the ratios at other stages of the reaction change with
competing reactions. 3. The ratio II:III and III:I cannot
be regarded as const. when the ratio of II:I is variable. 4.
The fact that addn. (75 min.) of BaNH₂ and III to the prod-
ucts at 80° did not lead to a decrease in the yields over the
control, causes V. to doubt the validity of the expts. II
is important in that it lowers the concn. of III and shifts the
equil. toward formation of IV. The yield of V is governed
by the compn. of the reaction mixt. at the moment of com-
pletion of the reaction if time is allowed for attainment of
equil.
Jane C. Aycock